Even More Practice Writing Formulas for Chemical Compounds

Part 1: Simple Binary Compounds

Give the formulas for the following binary compounds.

1. sodium fluoride
2. lithium sulfide
3. calcium bromide
4. strontium iodide
5. potassium oxide
6. magnesium nitride
7. silver iodide
8. cadmium selenide
9. aluminum sulfide
10. lithium bromide
11. sodium oxide
12. barium sulfide
13. aluminum nitride
14. beryllium oxide
15. zinc phosphide
16. potassium chloride
17. sulfur trioxide
18. phosphorus pentachloride
19. dinitrogen tetroxide
20. silicon dioxide

Part 2: Compounds with Polyatomic Ions

Write the correct formulas for each of these compounds.

1. sodium sulfite
2. aluminum phosphate
3. potassium bicarbonate (or potassium hydrogen carbonate)
4. beryllium sulfate
5. lithium carbonate
6. ammonium fluoride
7. sodium nitrate
8. potassium peroxide
9. calcium dichromate
10. magnesium sulfite
11. magnesium hyposulfite
12. strontium sulfate
13. aluminum hydroxide
14. lithium phosphate
15. sodium nitrite
16. calcium acetate
17. ammonium phosphate
18. barium cyanide
19. calcium bromite
20. cadmium hypochlorite

Part 3: Compounds with Type II Cations

Write formulas for the following compounds containing Type II (variable charge) cations.

1. iron (II) chloride
2. iron (III) chloride
3. chromium (III) oxide
4. mercury (II) sulfide
5. lead (IV) chlorate
6. cobalt (II) chlorite
7. chromium (VI) cyanide
8. tin (II) hydroxide
9. iron (III) dichromate
10. nickel (III) acetate